

3.1

p160 #3, 5, 7 p178 #16, 18, 20, 22, 23 p181-182 #27, 36, 38, 39, 43, 46 (not d)
49, 50, 61 (not d), 62

3. increasing atomic #

5. b

7. Li, K, Rb, Cs, or Fr

16. → dec ↓ inc

18. → inc ↓ dec

20. → inc ↓ dec

22. Na, Al, S, Cl ; period

23a. Na

b. P

27. yes same valence e^-

36a. Na

b. Sr

c. Ge

d. Se

38a. B

b. Mg

c. Al

39a. Sr, Mg, Be

b. Cs, Ba, Bi

c. Na, Al, S

43a. F

b. N

c. Mg

d. As

46. a, c

49. b same valence e^-

50. F has more p^+ than O so greater attraction; fewer e^- shells than Cl

61a. electronegativity increases across periods

62. zinc has a smaller radius so it has a stronger attraction to e^- so more energy is required to remove it