1. Define covalent bond. How can you tell if a bond will be covalent or ionic?
2. What is happening to the electrons in a covalent bond? How does that fulfill octet rule?
3. Why would a covalent bond form instead of an ionic bonds in terms of electronegativity?
4. What is a molecule?
5. How many electrons are involved in double bond? How many in triple bond?
6. List the single, double, and triple bonds in order of increasing strength. List the single, double, and triple bonds in order of increasing bond length.
7. What are the properties of a covalent bond/molecular compound?
8. A compound is a red crystalline solid that is brittle and conducts electricity in solution. Would this compound be classified as ionic or covalent?
9. Label the following compounds as ionic or covalent:
	1. CuCl2
	2. C6H6
	3. N2O5
	4. Ti(CN)4
10. Draw Lewis Structures for the following compounds:
	1. I2
	2. OF2
	3. SO3
	4. SOCl2 (Cl = chlorine)
	5. HF
	6. BO21-
	7. GeH2I2