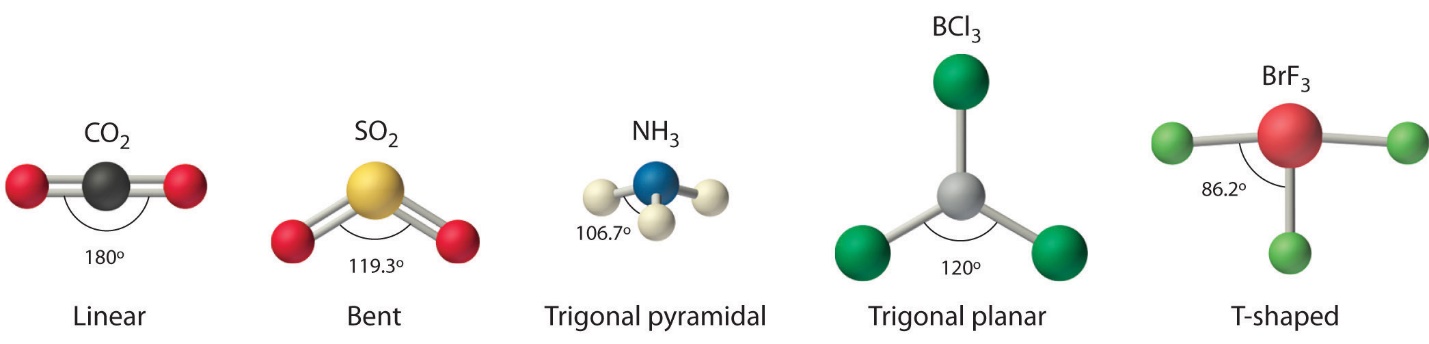
1. Describe VSEPR.
2. Using VSEPR explain why a molecule would take on a bent shape instead of a linear shape.
3. Answer the following questions using the picture below.

A B C D



* 1. Which molecule(s) have no lone pairs on the central atom?
  2. Which molecule is trigonal planar?

1. What type of shape consists of 2 atoms bonded to the central atom with 2 lone pairs?
2. Fill in the following chart

|  |  |  |
| --- | --- | --- |
| Molecule | Lewis Structure | Molecular geometry (Shape) |
| SiCl4  (Cl is chlorine) |  |  |
| CO32- |  |  |
| SCl2  (Cl is chlorine) |  |  |
| Molecule | Lewis Structure | Molecular geometry (Shape) |
| CO2 |  |  |
| CO |  | *Hint: What is the only shape it could look like?* |
| H2Se |  |  |
| SeO3 |  |  |
| H3O+ |  |  |
| PBr5 |  |  |
| RnF4 |  |  |
| TeF6 |  |  |