

Name:
Period:
Date:

6.5 Tracked Assignment (Obj 3)

Part 1: Answer the following questions.

1. What determines if one metal will replace another metal in a single replacement reaction?

if the element is more active or reactive than one in the compound they'll switch

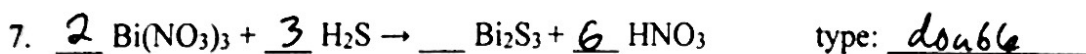
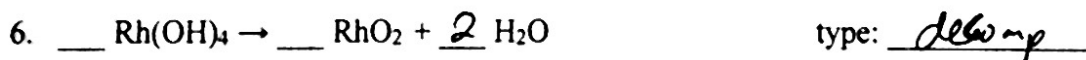
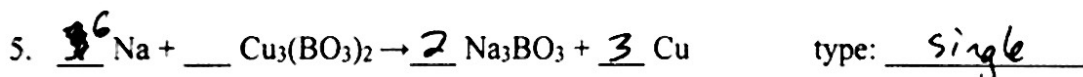
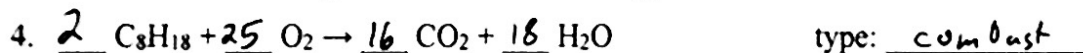
2. Label the following with the type of reaction expected:

- a. Two compounds reacting to produce two more *double*
b. Two elements combining *comb/syn*
c. An element and an ionic compound reacting *single*
d. The products are CO₂ and H₂O *combust*
e. One compound reacts *decomp*

3. What type of reaction is represented below? Write a balanced chemical equation for it. *single* $AB + C \rightarrow AC + B$

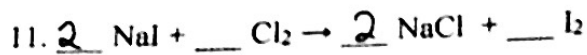


Part 2: Balance the following reactions and state the type of reaction involved.

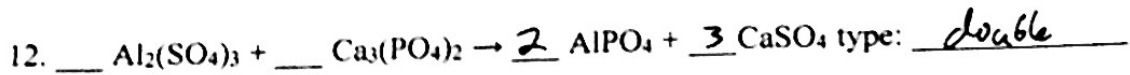


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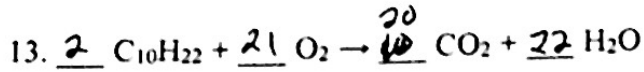
6.5 Tracked Assignment (Obj 3)



type: single

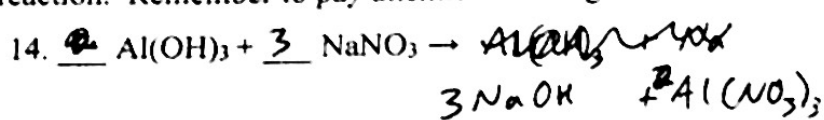


type: double

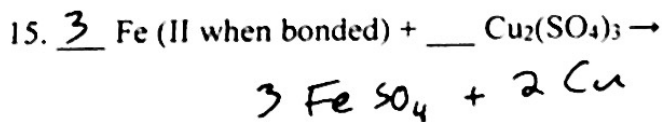


type: combust

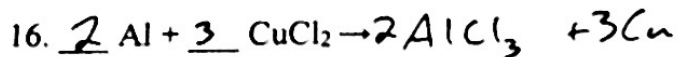
Part 3: Write a **complete balanced** equation for each of the reaction and **identify** the type of reaction. Remember to pay attention to charges and the activity series.



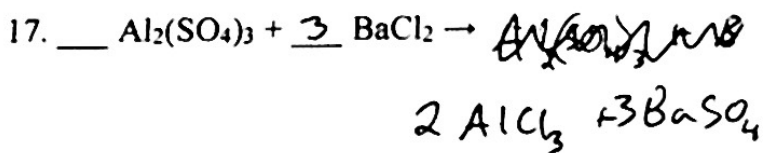
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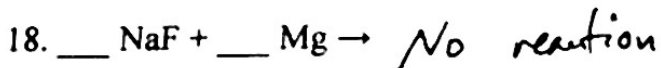
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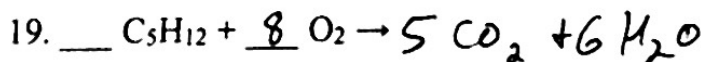
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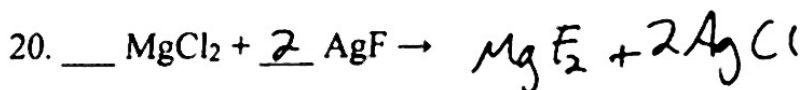
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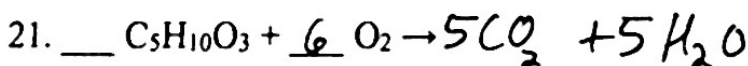
type: single



type: combust



type: double



type: Combust